

Department of Materials and Metallurgical Engineering
South Dakota School of Mines and Technology

Possible Blast Furnace Phosphorous Reactions						$\Delta G^{\circ} = A + BT$								
						A (Kcal)	- B (cal/K)	Temperature Range, °C						
1.	$\text{Ca}_3(\text{PO}_4)_2(\text{s})$	=	3.0	$\text{CaO}(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	553.4	140.0	25 –			
2.	$\text{Ca}_3(\text{PO}_4)_2(\text{s})$	+	3.0	$\text{SiO}_2(\text{s})$	=	3.0	$\text{CaSiO}_3(\text{s, L})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	489.5 493.7	139.6 142.5	25 – 1210 1210 – 1543
3.	$\text{Ca}_4\text{P}_2\text{O}_9(\text{s})$	=	4.0	$\text{CaO}(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	563.6	144.0	25 –			
4.	$\text{Ca}_4\text{P}_2\text{O}_9(\text{s})$	+	4.0	$\text{SiO}_2(\text{s})$	=	4.0	$\text{CaSiO}_3(\text{s, L})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	478.4 484.0	143.5 147.3	25 – 1210 1210 – 1543
5.	$\text{Fe}_3(\text{PO}_4)_2(\text{s})$	=	3.0	$\text{FeO}(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	480.5	155.8	25 – 1238			
6.	$\text{Fe}_3(\text{PO}_4)_2(\text{s, L})$	+	1.5	SiO_2	=	1.5	$\text{Fe}_2\text{SiO}_4(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	468.6 502.8	150.3 173.0	25 – 1217 1217 – 1238
7.	$\text{Mg}_3(\text{PO}_4)_2(\text{s})$	=	3.0	$\text{MgO}(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	501.6	144.0	25 –			
8.	$\text{Mg}_3(\text{PO}_4)_2(\text{s})$	+	3.0	$\text{SiO}_2(\text{s})$	=	3.0	$\text{MgSiO}_3(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	474.9	140.7	25 – 1300
9.	$\text{Mn}_3(\text{PO}_4)_2(\text{s})$	=	3.0	$\text{MnO}(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	495.9	156.0	25 – 1119			
10.	$\text{Mn}_3(\text{PO}_4)_2(\text{s})$	+	3.0	$\text{SiO}_2(\text{s})$	=	3.0	$\text{MnSiO}_3(\text{s})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	478.1	147.0	25 – 1119
11.	2 $\text{Na}_3\text{PO}_4(\text{s})$	=	6.0	$\text{Na}(\text{g})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	1106.2	366.3	925			
12.	2 $\text{Na}_3\text{PO}_4(\text{s})$	+	6.0	$\text{SiO}_2(\text{s})$	=	3.0	$\text{Na}_2\text{Si}_2\text{O}_5(\text{L})$	+	$\text{P}_2(\text{g})$	+	$2.5 \text{O}_2(\text{g})$	1779.0	1434.0	925
13.	2 $[\underline{\text{P}}]_{(1 \text{ wt\% in Fe (L)})}$	=					$\text{P}_2(\text{g})$		101.6	6.6				
14.	2 $\text{Fe}_3\text{P}(\text{s})$	=	6.0	$\text{Fe}(\text{s})$	+	$\text{P}_2(\text{g})$			102.0	22.6				
15.	2 $\text{Fe}_2\text{P}(\text{s})$	=	4.0	$\text{Fe}(\text{s})$	+	$\text{P}_2(\text{g})$			100.0	22.6				
16.	4 $\text{FeP}(\text{s})$	=	2.0	$\text{Fe}_2\text{P}(\text{s})$	+	$\text{P}_2(\text{g})$			37.8	22.0				
17.	2 $\text{FeP}_2(\text{s})$	=	2.0	$\text{FeP}(\text{s})$	+	$\text{P}_2(\text{g})$			58.7	43.0				

References

- K. L. Komarek, Trans TMS-AIME, 1963, vol. 227, pp. 136-145.
- U.S. Dept. of Commerce, MBS: JANAF Thermochemical Tables, 2 ed., U.S. Govt Printing Office, Washington, D.C., 1971